**Learning Goals:**

* Design experiments to measure the relationships between pressure, volume, and temperature.
* Create graphs based on predictions and observations.
* Make qualitative statements about the relationships between pressure, volume and temperature using molecular models.

**Predictions:** Make a chart like the one below. Without using the simulation, sketch what you think the graphs would look like. **Note: Be sure to label your x and y axes.**

|  |  |
| --- | --- |
| 1. Volume-Pressure graph

xy | Explain your reasoning for the graph’s appearance |
| 1. Volume-Temperature graph

xy | Explain your reasoning for the graph’s appearance |
| 1. Temperature-Pressure graph

xy | Explain your reasoning for the graph’s appearance |
| 1. Number of particles – Volume

xy | Explain your reasoning for the graph’s appearance |

**Experiments:**

1. For each case, I-IV. Write a short description of how to use the sim to collect data. Then collect data (table below) and make a graph for each set. Some helpful hints – if you set a parameter like temperature constant, then make a change, you may have to watch the temperature adjust and not record your data until the temperature is back to the original setting. These experiments would be difficult in a real situation because it is complicated to isolate parameters like you can in the sim.

|  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- |
| **Vol** | **Press** | **Vol** | **Temp** | **Temp** | **Press** | **Mole** | **Vol** |
|  |  |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |

(Roughly sketch your labeled graphs below)

1. After you have made your graphs (in Excel, on your calculator, wherever), check your predictions, and see if any might need some corrections. If necessary, make corrections in a different color including corrections to your reasoning.
2. Complete this table:

|  |  |  |  |
| --- | --- | --- | --- |
| Relationship | Direct or indirect? | Constant parameters | Whose Law? |
| V vs P |  |  |  |
| V vs T |  |  |  |
| T vs P |  |  |  |
| Moles vs V |  |  |  |

1. Using your results, explain each of the following scenarios. Make sure to refer to which graph can be used as evidence for your answer.
2. Explain why bicycle tires seem more flat in the winter than in summer.
3. Explain why a can of soda pop explodes if left in the hot sun.
4. A rigid container filled with a gas is placed in ice (ex. nalgene bottle). What will happen to the pressure of the gas? What do you think will happen to the volume?
5. An infected tooth forms an abscess (area of infected tissue) that fills with gas. The abscess puts pressure on the nerve of the tooth, causing a toothache. While waiting to see a dentist, the person with the toothache tried to relieve the pain by treating the infected area with moist heat. Will this treatment help? Why or why not?
6. Explain why adding air to a flat tire changes both the volume and the pressure.
7. What is the ‘particle model’ when it comes to gases?

|  |  |  |  |
| --- | --- | --- | --- |
| Relationship | Direct or indirect? | Constant parameters | Briefly, why according to particle model. |
| V vs P |  |  |  |
| V vs T |  |  |  |
| T vs P |  |  |  |
| Moles vs V |  |  |  |

1. Fill in the table below with your particle model explanation for each interaction.
2. Summarize your explanations from the table above into a single statement.
3. Create a single proportionality equation demonstrating how volume is related to moles, pressure, and temperature. (i.e., volume is on one side, moles/pressure/temperature are on the other, and there’s no equals sign used.)
4. Do the gas laws have any limitations? Explain.